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25.2 g of $\text{CuSO}_4 \cdot 6\text{H}_2\text{O}$ is dissolved in 28.0 mL of water, calculate the molarity.
 $25.2 \text{ g} \times \frac{1 \text{ mole}}{267.72 \text{ g}} = 3.36 \text{ M}$
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Write your answer. Round off the number of digits after the decimal point to two or three, depending on your instructor's preference. When you write out the answer, abbreviate "molarity" with "M" and state the chemical abbreviation of the solute involved. Example problem: 0.179 M NaCl

4 Ways to Calculate Molarity - wikiHow

Problem #1: Sea water contains roughly 28.0 g of NaCl per liter. What is the molarity of sodium chloride in sea water? Solution: $MV = \text{grams} / \text{molar mass} (x) (1.00 \text{ L}) = 28.0 \text{ g} / 58.443 \text{ g mol}^{-1}$. $x = 0.4790993 \text{ M}$ to three significant figures, 0.479 M

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Solved: (18pts)4. What Would Be The Molarity 28.24 Ml Of S

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Calculate the molarity of 28.9 g of MgS in 843 mL of solution.? Answer Save. 1 Answer. Relevance. ChemTeam. Lv 7. = mass / molar mass (x) (0.843 L) = 28.9 g / 56.37 g/mol. x = 0.608 M (to three sig figs) 0 0. Still have questions? Get your answers by asking now. Ask Question + 100. Join Yahoo Answers and get 100 points today. Join. Trending

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Steps to find answer: 1. Calculate molecular weight of MgS. 2. Calculate Molarity. 26.5 g x 1 Mole/molecular wt of MgS(g) x1/0.897 L. The resulting answer once all the units are canceled out is moles /liter or Molar

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What is the molarity? 214.2g OsF 3 x 1 mol OsF 3 = 12.9 M OsF 3. 0.0673 L soln 247.23 g OsF 3. Calculate the molarity if a flask contains 1.54 moles potassium sulfate in 125 ml of solution. 1.54 mol K 2 SO 4 = 12.3 M K 2 SO 4. 0.125 L soln. A chalice contains 36.45 grams ammonium chlorite in 2.36 liters of solution - calculate the molarity.

What is the molarity of 26.5 g of MgS in - Yahoo Answers

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(18pts)4. What would be the molarity 28.24 ml of solution of a phosphoric acid used in neutralize 151.54 g of Ruthenium Hydroxide ($\text{Ru}(\text{OH})_2$):. $2 \text{HPO}_4 + 3 \text{Ru}(\text{OH})_2 \rightarrow \text{Ru}_3(\text{PO}_4)_2 + 6 \text{H}_2\text{O}$ + (18pts). A solution Manganese nitrate [$\text{Mn}(\text{NO}_3)_2$] solution is made in which 188.0g solute is dissolved in mass of water was 752.0g of water.

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The molarity of a solution is calculated by taking the moles of solute and dividing by the liters of solution. This is probably easiest to explain with examples. Example #1: Suppose we had 1.00 mole of sucrose (its mass is about 342.3 grams) and proceeded to mix it into some water. It would dissolve and make sugar water.

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The answer represents the moles of glucose per kilogram of water and has three significant figures. Molality and molarity are closely related in value for dilute aqueous solutions because the density of those solutions is relatively close to 1.0 g/mL. This means that 1.0 L of solution has nearly a mass of 1.0 kg.

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